

CHAPTER 4

Increase in Metal Extractability After Liming of Sacrificial Sewage Sludge Disposal Soils

4.1 Abstract

A sandy and a sandy clay-loam soil from two dedicated (sacrificial) sewage sludge disposal sites were incubated with a total lime equivalent of $45 \text{ Mg}\cdot\text{ha}^{-1}$. Both these soils were acidified (pH 4.0 to 4.2) and had a significant accumulation of organic material (organic C of 2.9 to 3.7 %) compared to non-polluted soils (organic C of 0.6 to 1.0 %). The limed soils did not attain the desired pH of 6.5 after 6 months' incubation due to a high buffer capacity. After incubation soil samples were taken from the incubated pots and the levels of Al, Fe, Mn, Cu, Zn, Pb and Cd were determined by atomic absorption spectrophotometry (AAS) after extraction with NH_4 -EDTA and BaCl_2 . Most of the metals extracted with BaCl_2 (except Mn in the sandy soil and Cd in both soils) decreased after liming. The EDTA-extractable Mn, Fe, and Cd in both soils and Cu and Pb in the sandy clay-loam soil increased after liming, whereas Al and Zn decreased in extractability. (Statistically significant differences could not be determined for the trial due the trial not having been designed for the results that were obtained.) Similar results were reported in literature for EDTA metal extraction but the phenomenon was not elaborated upon, except for Cr. The increased extractability of some of the metals after liming could negatively influence the use of EDTA as an extracting agent in proposed heavy-metal guidelines for similar sacrificial soils. Should liming be considered as a strategy to decrease metal mobility in sacrificial soils, the observed increase in extractability becomes a cause for concern and should receive attention in further research.

4.2 Introduction

In recent years the trend has been to express heavy-metal levels in soils as plant- or potentially plant-available metal levels (Beckett, 1989; McLaughlin et al., 2000) rather than "total" concentrations. This implies that "weaker" extractants or chelating agents be used in heavy-metal studies. Ethylenediaminetetraaceticacid (EDTA), in

either the di-sodium or di-ammonium salt form, has been used extensively in a host of studies as an extractant of potentially plant-available heavy metals.

In some trials, EDTA was found to give a very good indication of the pollution hazard of heavy metals in soils as well as being a reliable test for predicting plant-available metals (Hooda et al., 1997; Cajuste and Laird, 2000). Earlier, Bruemmer and Van der Merwe (1989) stated that the NH_4 -EDTA-extractable heavy metal concentration gives a good estimate of those potentially plant-available, and therefore suggested it to be used in the establishment of preliminary threshold values for heavy metals in South African soils. Currently, however, there are no guidelines stipulating the maximum EDTA-extractable metal levels in South African soils.

Neutral salt extractants are generally weaker extractants than EDTA and give an indication of the immediately exchangeable (therefore immediately plant-available) metals (Beckett, 1989; McLaughlin et al., 2000). Examples of such extractants are BaCl_2 , NH_4NO_3 , NH_4OAc buffered at pH 7, and more. The BaCl_2 method (Hendershot and Duquette, 1986) gives an indication of the effective cation exchange capacity (ECEC) of the soil at unbuffered pH levels. This is particularly relevant in studies where the pH dependence of metal extractability is one of the parameters of investigation.

In a preliminary study it was found that soils from some sacrificial sewage sludge disposal sites can be acidified (pH 4.0 to 4.2) and also have very high pH-buffering capacities due to relatively high organic carbon levels (organic C of 2.9 to 3.7 %). Coupled to this is a significant increase in total heavy metal content of the soil. The aim of this study was to determine the BaCl_2 and NH_4 -EDTA extractability of a range of metals in two acid soils after liming to near-neutral pH levels and incubation in pots.

4.3 Materials and Methods

The buffer capacity of the two acid soils (here referred to as Soils 1 and 2) was determined with a $\text{Ca}(\text{OH})_2$ buffer (as discussed by Van der Waals and Claassens, 2002). The required amount of a commercial dolomitic lime (according to the buffer

determination) for a pH of 6.5 was added to each soil in 7.5kg pots (with 4 repetitions) and the soil incubated for 3 months with regular watering and mixing. After sampling the soil and finding only a slight change in pH it was decided to add an equal amount of lime and incubate the soil again for the same period of time. The total amount of lime added amounted to the equivalent of $45 \text{ Mg}\cdot\text{ha}^{-1}$.

After the second incubation period a representative sample was taken from each pot and the pH determined according to the method described by the The Non-Affiliated Soil Analysis Work Committee (1990). On the limed soil and a sample from the original soil a BaCl_2 extraction was done. The method was adapted from Hendershot and Duquette (1986) by shaking 5 g of soil with 50 ml 0.1 M BaCl_2 in a glass bottle on a horizontal shaker for 1 h. After filtering the solution, the Al, Mn, Fe, Cu, Zn, Pb, and Cd contents were determined through AAS. Further, an EDTA extraction (The Non-Affiliated Soil Analysis Work Committee, 1990) was done on the same samples and Al, Fe, Mn, Cu, Zn, Pb and Cd again determined by AAS. The samples were tested at the same time to minimise experimental error differences. The metals Cr and Ni were not determined due to the termination of the testing after the first results had yielded unexpected increases upon liming, as discussed below.

4.4 Results and Discussion

Table 4.1 gives the pH results of the two soils before and after liming. The pH values increased by 1.4 pH units for Soil 1 and by 1.8 pH units for Soil 2 after the addition of the equivalent of $45 \text{ Mg}\cdot\text{ha}^{-1}$ lime. This indicates a massive buffering capacity brought about by organic material and complexes stable at pH 4.2 and 4.0 for Soil 1 and 2 respectively. The buffer determination as described by Van der Waals and Claassens (2002) is therefore considered inadequate for soils with high buffer capacities such as these.

Table 4.1. Soil 1 and 2 pH values before and after liming (n = 4)

Method	Soil 1		Soil 2	
	Before	After	Before	After
Water	4.2	5.6	4.0	5.8
CaCl ₂	3.8	5.3	3.7	5.7
KCl	3.3	5.0	3.5	5.4

Table 4.2 indicates the influence of the lime addition and incubation on the BaCl₂ extractable metals from the two soils. The metals Mn and Cd in Soil 1 and Cd in Soil 2 did not decrease in extractability after liming as was expected. The expected decrease was found for Al, Cu, Fe, and Zn. The large coefficients of variation for the limed soils are indicative of the low values determined for some of the metals that approached their detection limit with AAS.

Table 4.3 indicates the EDTA-extractable metals for the two soils before and after liming. Here most of the tested metals (Cu, Mn, Fe, and Cd in both soils and Pb in Soil 2) did not decrease in extractability and in some cases even increased. Aluminium, Zn, and to a lesser extent Pb in Soil 1, indicated the expected decrease in extractability.

In a study by Bloomfield and Pruden (1975) an increase in the EDTA extractability of Cd, Cu, Cr, Ni, Pb, and Zn in limed sludge samples and an increased extractability of Cr in limed “sludge + soil samples” was reported. Lake et al. (1984) attributed this increase in Cr levels to the chemistry of the Cr(VI) form in soils. In the Bloomfield and Pruden (1975) study no explanation was given for the increased extractability of metals other than Cr from the limed sludge samples. Although Zn decreased in extractability in the present study, the effect of increased extraction was still observed for Cu, Mn, Fe, Pb, and Cd, even though the samples are not exactly of the same type as those in the Bloomfield and Pruden (1975) study.

Table 4.2. The effect of liming on BaCl₂ extractable metals (mg·kg⁻¹ soil; n = 4; values in brackets denote the standard deviation)

Metal (mg·kg ⁻¹)	Soil 1		Soil 2	
	Before	After	Before	After
Al	35.0 (1.3)	0.2 (0.3)	154.7 (6.7)	0.2 (0.1)
Cu	3.3 (0.2)	0.2 (0.1)	14.1 (0.2)	0.8 (0.1)
Mn	17.5 (0.5)	21.2 (4.0)	29.7 (2.0)	12.4 (3.0)
Fe	6.0 (1.0)	0.5 (0.5)	0.8 (0.8)	0.3 (0.2)
Zn	151.6 (6.4)	11.5 (3.2)	18.5 (0.3)	3.0 (0.4)
Pb	1.8 (2.7)	1.3 (0.7)	1.6 (1.9)	0.3 (0.3)
Cd	0.3 (0.2)	0.3 (0.1)	0.3 (0.2)	0.6 (0.1)
Coefficient of Variation	5.7	25.4	5.6	24.4

Table 4.3. Effect of liming on NH₄-EDTA-extractable metals (mg·kg⁻¹ soil; n = 4; values in brackets denote the standard deviation)

Metal (mg·kg ⁻¹)	Soil 1		Soil 2	
	Before	After	Before	After
Al	113.9 (10.4)	33.2 (3.8)	337.7 (16.5)	222.2 (55.6)
Cu	41.1 (1.8)	37.1 (2.2)	80.8 (1.4)	104.6 (2.9)
Mn	17.2 (0.6)	49.6 (3.8)	37.1 (5.9)	71.3 (6.9)
Fe	242.7 (2.7)	383.5 (11.9)	170.7 (5.6)	271.7 (23.9)
Zn	117.0 (8.2)	49.1 (2.0)	28.8 (1.5)	15.6 (0.8)
Pb	9.8 (0.7)	6.5 (0.3)	8.4 (2.6)	8.1 (1.0)
Cd	0.3 (0.02)	0.8 (0.2)	0.5 (0.01)	1.3 (0.1)
Coefficient of Variation	4.5	4.3	5.1	13.1

4.5 Conclusions and recommendations

The results in this trial lead to a number of questions concerning the extractability of the metals and the conditions in the pots. Firstly, although the lime used was a commercially available lime, the metals listed here did not occur at sufficiently high levels to lead to the resultant increase in extractability (data not presented here). It is also not clear whether the lime had reacted completely with the soil. Secondly, the possible fluctuation in pH during the incubation time is not known and it is therefore impossible to comment on its influence on the metal extractability. Thirdly, the possible mineralization of organic material after liming could not be quantified due to the set-up of the trial (aerobic conditions in a greenhouse) and it is therefore not possible to comment on its influence on the increased extractability.

Although results similar to these have not been widely reported, the phenomenon warrants further investigation. These authors agree with the conclusions of Lake et al. (1984) concerning the increased Cr extractability in the Bloomfield and Pruden (1975) study, but consider the phenomenon of the increased extractability of the metals other than Cr as similar to those reported here. The metals that exhibited an increase in extractability do not occur in soil in the same state as Cr(VI) and therefore require a different explanation.

The trial, having generated a number of questions, was not designed to supply the required answers. It is therefore suggested that a dedicated trial be conducted to determine:

- The change in pH over an extended incubation period with and without the addition of lime with regular sampling intervals.
- The change in metal extractability (complexed and exchangeable) over an adequate incubation period with and without the addition of lime with regular sampling intervals.
- The EDTA metal extractability as influenced by increasing pH (lime application rates).

The increased extractability of the metals after the application of lime has a profound influence on the establishment of metal guidelines when EDTA is used. Before the use of EDTA as an extractant in guideline levels is advocated, this phenomenon has to be investigated further to determine the restrictions of the procedure – especially under the conditions experienced in this trial.