

Heterogeneous electron transfer kinetics and electrocatalytic behaviour of mixed self-assembled ferrocenes and SWCNT layers

Dudzile Nkosi,^a Jeseelan Pillay,^a Kenneth I. Ozoemena,^{*ab} Khalid Nouneh^{cd} and Munetaka Oyama^c

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The electron transfer dynamics and electrocatalytic behaviour of ferrocene-terminated self-assembled monolayers (SAMs), co-adsorbed with single-walled carbon nanotubes (SWCNTs) on a gold electrode, have been interrogated for the first time. Ferrocene monocarboxylic acid (FMCA) or ferrocene dicarboxylic acid (FDCA) was covalently attached to the cysteamine (Cys) monolayer to form Au-Cys-FMCA and Au-Cys-FDCA, respectively. The same covalent attachment strategy was used to form the mixed SWCNTs and ferrocene-terminated layers (*i.e.* Au-Cys-SWCNT/FMCA and Au-Cys-SWCNT/FDCA). Using cyclic voltammetry and electrochemical impedance spectroscopy, the impact of neighbouring SWCNTs on the electron transfer dynamics of the ferrocene molecular assemblies in an acidic medium (0.5 M H₂SO₄) and in a solution of an outer-sphere redox probe ([Fe(CN)₆]⁴⁻/[Fe(CN)₆]³⁻) was explored. The electron transfer rate constants in both media essentially decreased as Au-Cys-FMCA > Au-Cys-SWCNT/FDCA > Au-Cys-FDCA > Au-Cys-SWCNT/FMCA. This trend has been interpreted in terms of several factors such as the locations of the ferrocene species in a range of environments with a range of potentials, the proximity/interactions of the ferrocenes with one another, and electrostatic interaction or repulsion existing between the negatively-charged redox probe and the modified electrodes. The thiocyanate ion (SCN⁻) was used as a model analyte to examine the influence of the neighbouring SWCNTs on the electrocatalytic ability of the ferrocene assemblies. The Au-Cys-SWCNT/FDCA showed the best catalytic activity (in terms of onset potential and catalytic peak current height) for the oxidation of SCN⁻, possibly due to the repulsive interactions between the negatively charged SCN⁻ and high number of surface -COOH species at the SWCNT/FDCA. This study has provided some useful insights as to how CNTs co-assembled with ferrocene-terminated thiols could impact on the electron transfer kinetics as well as the electrocatalytic detection of the self-assembled ferrocene layers.

1. Introduction

Ferrocene (Fc) and its derivatives are well known electrocatalysts. One of the interesting techniques of immobilising them onto electrode surface is by self-assembly. Self-assembly simply describes the science of things that put themselves together.¹ A self-assembled monolayer (SAM) has been described as an ideal system for disentangling the fundamental events in interfacial electron transfer.² Several reports have established that the redox kinetics of self-assembled monolayers of ferrocenes are influenced by other molecules that share the same local microenvironments with them. For example, Weber *et al.*³ reported that electron transfer (ET) is dependent on the chain length of the mercapto-alkanol to

which the Fc is attached. Sumner and Creager⁴ showed that the electron transport of Fc SAM is related to the extent to which alkanethiols permit the exposure of the Fc on to the electrolyte. Recently, Hortholary and Crouchet⁵ showed that an Ru complex co-assembled with Fc on gold exerts some influence on the heterogeneous electron transfer (HET) of the Fc. In addition, several workers have also shown that thiol adsorbates containing ferrocene head groups are important for interrogating the chemical features of monolayers on gold substrates. For example, Collman *et al.*⁶ used ferrocene adsorption to demonstrate the applicability of Sharpless "click" chemistry as a general methodology for functionalizing surfaces coated with self-assembled monolayers. Chidsey and co-workers² studied co-adsorbed ferrocene and unsubstituted alkanethiol and observed the structural orientation/changes that accompany concentration changes; Aluletta *et al.*⁷ studied mixed SAMs of ferrocene-terminated thiols and hydroxy-alkanethiols and proved the occurrence of phase separations; Ye *et al.*⁸ observed coverage-dependent behavior of the ferrocene-terminated thiol SAMs induced by the redox reaction of the terminal ferrocene moiety, while Waldeck and co-workers⁹ studied ferrocene co-assembled with alkanethiolate SAMs and

^a Department of Chemistry, University of Pretoria, South Africa

^b Energy & Processes Unit, Materials Science and Manufacturing, Council for Scientific and Industrial Research (CSIR), Pretoria 0001, South Africa. E-mail: kozoemena@csir.co.za; Fax: +27 (0)12 841 2135; Tel: +27 (0)12 841 3664

^c Department of Material Chemistry, Graduate School of Engineering, Kyoto University, Nishikyo-ku, Kyoto 615-8520, Japan

^d The Institute of Nanomaterials and Nanotechnology (INanotech), Rabat, Morocco

showed how intermolecular electronic coupling pathways could contribute to the electron transfer; Uosaki and co-workers¹⁰ studied SAM containing both ferrocene and azobenzene moieties and established how electro- and photochemical structural conversion of the *cis* and *trans* forms of the azobenzene moiety could impact on the charge-transfer processes; Crooks and co-workers¹¹ studied mixed monolayers of ferrocene-dendrimers and alkanethiols and, for the first time, introduced an elegant strategy for constructing an electrochemical current rectifier that permits current flow in only one direction; and Chambers *et al.*¹² reported how a ferrocene-thiol SAM co-adsorbed with alkanethiolates could unravel nanoscale phase separation.

To our knowledge, there is no report on ferrocene-terminated SAMs co-adsorbed with CNTs. Hence, until now, it was not known how CNTs co-assembled with ferrocene-terminated thiols could impact on the electrochemistry of ferrocene moiety. The only known reports on CNTs are those where Fcs were covalently attached onto the ends of CNT SAMs *via* amide^{13,14} or ester¹⁵ bonds. This somewhat surprising considering that CNTs could enhance intermolecular electron transfer between ferrocene sites by behaving as electronic bridges. To address this discrepancy, we show for the first time, the electrochemistry of mixed SAMs of single-walled carbon nanotubes (SWCNTs) and Fcs and examined the influence of SWCNTs on the HET of immobilized Fcs and electrocatalytic detection of a thiocyanate as an analytical probe. Thiocyanate is well known for its environmental and clinical importance.^{16–18} We show that CNTs have distinct influence on the HET and electrocatalytic properties of these self-assembled Fc complexes. In summary, this study is crucial as it could provide some useful physical insights into the impact of local microenvironment surrounding a redox-active molecule (not only ferrocene but other organometallic species) on the dynamics of charge transport between a redox molecule and electrode. Such insights will create some knowledge on the factors that must be considered when designing CNT-based molecular-scale electronics or electrocatalytic devices that transport charges within organized molecular assemblies in such a manner as to accomplish specific functions.

2. Experimental

2.1 Materials and methods

Ferrocenemonocarboxylic acid (FMCA) was from Fluka. Ferrocenedicarboxylic acid (FDCA), cysteamine hydrochloride (Cys), dicyclohexylcarbodiimide (DCC) and single-walled carbon nanotubes (SWCNTs) and *N,N*-dimethylformamide (DMF) were obtained from Sigma-Aldrich. The pristine SWCNTs were acid-functionalized as previously reported.^{19,20} Ultra pure water of resistivity 18.2 M Ω cm from a Milli-Q Water System (Millipore Corp., Bedford, MA, USA) was used throughout for the preparation of solutions. Phosphate buffered solutions (PBS) was prepared with NaH₂PO₄·H₂O. All other reagents were of analytical grades and were used as received from the suppliers without further purification.

AFM images were acquired with AFM 5100 System (Agilent Technologies, USA) using AC mode AFM scanner

interfaced with a PicoScan 5.0 controller. Silicon type PPP-NCH-20 (Nanosensors[®]) of thickness 4.0 ± 1.0 μm , length 125 ± 10 μm , width 30 ± 7.5 μm , spring constants $10\text{--}130$ N m⁻¹, resonant frequencies of 204–497 kHz and tip height of 10–15 μm were used. All images were taken in air at room temperature.

Electrochemical experiments were carried out using an Autolab Potentiostat PGSTAT 302 (Eco Chemie, Utrecht, The Netherlands) driven by the General Purpose Electrochemical Systems data processing software (GPES, software version 4.9). Electrochemical impedance spectroscopy (EIS) measurements were carried out with an Autolab FRA software between 0.1 Hz and 10 kHz using a 5 mV rms sinusoidal modulation. The FRA software allowed the fitting of the raw EIS data to equivalent circuit models using a *complex non-linear least squares* (CNLS) routine, with *Kramers–Kronig rule check*. The working electrode was bare gold ($r = 0.8$ mm, BAS) or the same gold electrode modified with the investigated SAMs. An Ag|AgCl wire and platinum wire were used as pseudo-reference and counter electrodes, respectively. As in a previous report,²¹ we used the reversible electrochemistry of [Fe(CN)₆]^{3-/4-}, assuming a diffusion coefficient of $\approx 6.3 \times 10^{-6}$ cm² s⁻¹,²² at scan rates ranging from 10 to 200 mV s⁻¹ and employing the Randles–Sevcik theory,²³ to estimate the electrochemical roughness factor of the gold electrode (*i.e.* ratio of the real (0.027 cm²) to geometrical (0.020 cm²) areas) and found it to be *ca.* 1.35. Silver/silver chloride coated with a porous layer of KCl and platinum were used as reference and counter electrodes, respectively. Prior to modification, the bare Au was first cleaned by first polishing in an aqueous slurry of alumina (<10 μm) on a SiC-emery paper (type 2400 grit), and then to a mirror finish on a Buehler felt pad. The electrode was then subjected to ultrasonic vibration in absolute ethanol to remove residual alumina particles that might be trapped at the surface. Finally the electrode was etched for about 2 min in a hot ‘piranha’ solution (1 : 3 (v/v) 30% H₂O₂ and concentrated H₂SO₄) and then rinsed with copious amounts of ultrapure Millipore water followed by ethanol. (*CAUTION: Piranha solution must be handled with care as it reacts violently with organic materials and can explode when stored in closed containers.*) This stage was necessary to remove organic contaminants. The cleanliness of the bare electrode surface was finally established by placing it in 0.5 M H₂SO₄ and scanning the potential between -0.5 and 1.5 V (*vs.* Ag|AgCl wire) at a scan rate of 0.05 Vs⁻¹ until a reproducible scan was obtained. After this check, the electrode was again rinsed with absolute ethanol and immediately placed into a nitrogen-saturated absolute ethanol solution of Cys (50 mg/120 ml) for 18 h at ambient temperature to form the Au-Cys.

2.2 Modification by sequential self-assembly on gold surface

Self-assembled monolayers of alkanethiol capped with carboxylated ferrocene and single-walled carbon nanotubes were prepared on gold surfaces. The gold electrode modified with cysteamine is represented as Au-Cys. The formation of Au-Cys-FDCA and Au-Cys-FMCA followed after placing the Au-Cys electrode in a 1 mL DMF solution containing FDCA

(1 mg, *ca.* 4 μmol) or FMCA (1 mg, *ca.* 4 μmol) and 0.5 mg DCC (2.4 μmol) for 48 h, respectively. The Au-Cys-SWCNT/FDCA electrode was obtained by placing Au-Cys in 1 mL DMF solution containing a mixture of SWCNT (~ 1 mg) and FDCA (1 mg, *ca.* 4 μmol) and 0.5 mg (2.4 μmol) DCC for 48 h. This same method was used to prepare Au-Cys-SWCNT/FMCA, using a mixture of ~ 1 mg each of SWCNT and FMCA dissolved in 1 mL DMF. This ratio is adopted throughout the study since it produced the best electrochemistry of the SAMs. Upon removal from the deposition solution, prior to electrochemical experiments, the electrode was thoroughly rinsed with millipore water and dried in a nitrogen atmosphere. The modified electrode was conditioned by placing in 0.5 M H_2SO_4 and repetitively scanned between -0.5 and 0.6 V (*vs.* Ag|AgCl) potential window at a scan rate of 25 mV s^{-1} until a constant scan was obtained.

3. Results and discussion

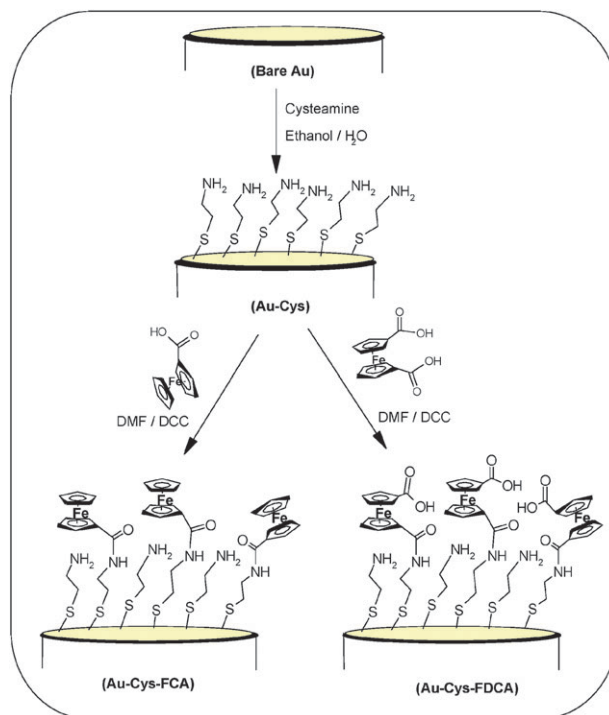
3.1 Electrode fabrication and AFM characterization

Scheme 1 summarizes the self-assembly process of the FMCA and FDCA onto the gold electrode using cysteamine as the base monolayer. Scheme 2, on the other hand, is a summary of the co-assembly of the SWCNTs and FMCA or FDCA. Both processes (Schemes 1 and 2) employed the normal carbodiimide coupling chemistry,^{20,24} whereby the $-\text{COOH}$ functional groups are first activated to permit the covalent bonding with the $-\text{NH}_2$ group of the cysteamine. Confirmation of the attachment of these redox-active species (FMCA, FDCA and SWCNTs) by sequential self-assembly technique was achieved, first, by the use of atomic force microscopy.

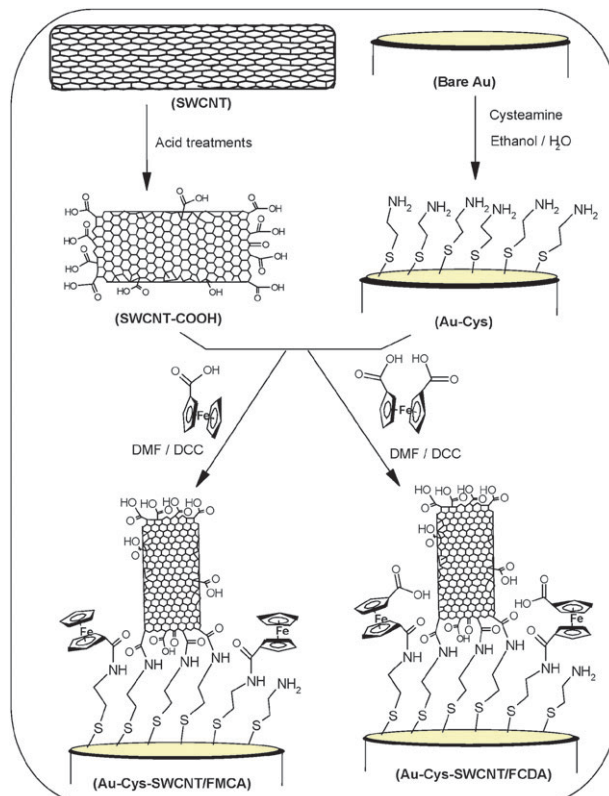
Fig. 1 is a comparison of the AFM images of (a) bare Au, (b) Au-Cys-SWCNT, (c) Au-Cys-FDCA and (d) Au-Cys-SWCNT/FDCA. The needle-like protrusion of the Au-Cys-SWCNT agrees with several reports for SWCNT-based SAMs.^{20,25–27} Typical lengths of acid-treated SWCNTs using the conditions used in this work lie in the 250–350 nm range.¹⁴ However, when we immobilized them as SAMs, the heights lie in the 30–50 nm range, which is in close agreement with literature.^{20,25} The SWCNTs assembled as bundles and not as individual tubes due to the strong van der Waals attractive forces existing between carbon nanotubes. The same explanation may be true for the needle-like protrusion shown by the Au-Cys-FDCA. Interestingly, as would be expected, the mixed layers, Au-Cys-SWCNT/FDCA, showed more compactness than either the individual layers. Generally, the AFM topographic height and roughness in Fig. 1 are greater with the combined SWCNT and Fc assemblies, exemplified with Au-Cys-SWCNT/FDCA (d). The roughness and height of the Au-Cys-SWCNT (b) are higher than those of the Au-Cys-FDCA (c) or Au-Cys-FMCA (not shown), confirming that SWCNTs are longer than the Fc assemblies (as indicated in the cartoon of Scheme 2).

3.2 Heterogeneous electron transfer dynamics in 0.5 M H_2SO_4 solution

The cyclic voltammetric properties of the ferrocene-modified gold electrodes were studied in 0.5 M H_2SO_4 solution. Fig. 2 shows the CV profiles of the electrodes obtained at 25 mV s^{-1} .



Scheme 1 Schematic of the gold electrode modification routes for Au-Cys, Au-Cys-FCA and Au-Cys-FDCA.



Scheme 2 Schematic of the gold electrode modification routes for Au-Cys-SWCNT/FMCA and Au-Cys-SWCNT/FDCA.

Each of the redox couples is ascribed to the Fe(III)/Fe(II) redox process. The electrochemical parameters, formal potential

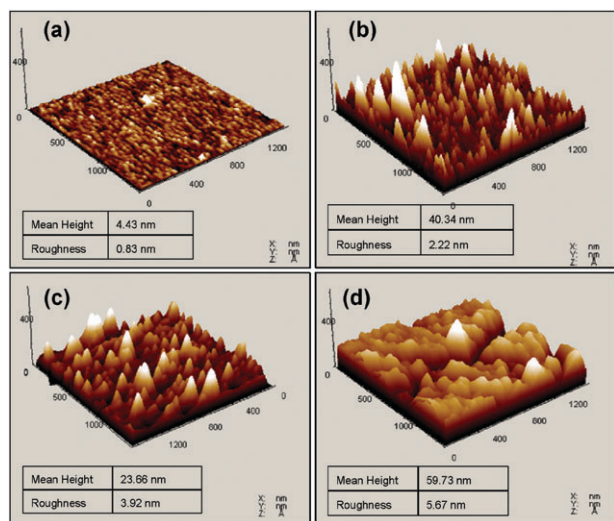


Fig. 1 Typical topographic images of the electrodes: (a) Au bare, (b) Au-Cys-SWCNT, (c) Au-Cys-FDCA and (d) Au-Cys-SWCNT/FDCA.

($E_{1/2}$ /mV), peak-to-peak potential separation (ΔE_p /mV), the width at half the peak current (E_{fwhm} /mV) and ratio of anodic peak current to the cathodic peak (I_{pa}/I_{pc}) and surface coverages (Γ /mol cm^{-2}) are summarized in Table 1. All the electrodes show $I_{pa}/I_{pc} \approx 1$, indicating voltammetric reversibility. The choice of the potential used here was based on our experience working with thiolated SAMs. As we cautioned in previous reports on SAMs,^{28,29} the thiol monolayer only show remarkable stability in acidic and slightly alkaline pH (pH 2–9) at potential between window of -0.5 and $+0.9$ V vs. Ag/AgCl. At more positive potential ($\geq +0.9$ V vs. Ag/AgCl) thiol desorption and/or Au surface oxidation could be observed. The stability observed for these SAMs may be related to some extent by hydrogen bonds arising from the $-\text{COOH}$ and $-\text{OH}$ functional groups. It may also be related to the protection of the sulfur by the large SWCNTs and and/or Fc rings. The interchain attractive interactions resulting from the alkyl chains may also be another contributing factor.

As seen from Table 1, the electron transport (signified by the magnitude ΔE_p) in this acid electrolyte condition is fastest at the Au-Cys-SWCNT/FDCA \approx Au-Cys-FMCA and slowest at the Au-Cys-SWCNT/FMCA. In a similar manner, the E_{fwhm} values (in the 71–117 mV range) slightly deviate from the ideal value of $90.6/n$ mV for $n = 1$,^{30–32} with the Au-Cys-SWCNT/FMCA (117 mV) being slightly greater than that of the Au-Cys-SWCNT/FDCA (104 mV). The higher values of ΔE_p and E_{fwhm} for the SWCNT/FMCA compared to its SWCNT/FDCA counterpart may be related to different factors. First, some workers^{13–15,33} have attributed this phenomenon to ferrocene species being located in a range of environments with a range of $E_{1/2}$. This should not be totally surprising if one considers the distribution of lengths of the SWCNTs on the gold substrate, and that the ferrocenes studied in this work could possibly be attached to the ends and/or defect sites of the SWCNTs *via* ester bonds (*i.e.* between the $-\text{COOH}$ of the

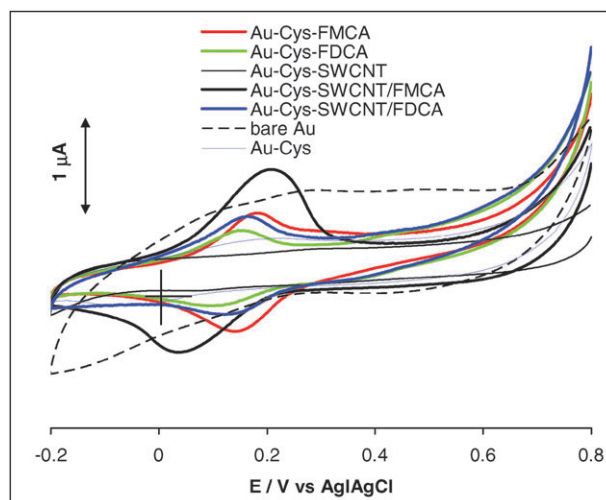


Fig. 2 Comparative cyclic voltammetric evolutions of the modified gold electrode in 0.5 M H_2SO_4 , Au-Cys-FMCA, Au-Cys-FDCA, Au-Cys-SWCNT, Au-Cys-SWCNT/FMCA and Au-Cys-SWCNT/FDCA. Scan rate = 25 mV s^{-1} .

ferrocene and the few $-\text{OH}$ groups on the SWCNTs). By different environments, it is meant that the ferrocene species have different formal potentials, and thus the effective voltammetric wave will consist of a superposition of distinct electrochemical responses, resulting in the observed voltammetric broadening.¹⁴ It is worth pointing out that the formal redox potentials of the two Fc species will be different. Also, as can be clearly seen from the CVs, especially in the 0.3–0.8 V region, the background current or capacitance of the bare Au (C_{Au}) is about three times higher than the capacitance of any of the SAM-modified gold electrode ($C_{\text{Au-SAM}}$). This is expected since in aqueous solution, unlike in organic solutions, the overall electrode double-layer capacitance should be governed by the bare Au (*i.e.* $C_{\text{Au}} > C_{\text{Au-SAM}}$).^{34,35}

The second explanation for the non-ideal ΔE_p and E_{fwhm} values for the adsorbed ferrocenes may be found from the works of Chidsey and co-workers.^{2,36,37} According to these workers, when the self-assembled ferrocene species are well separated and do not interact with one another, a narrow symmetric redox peak with ideal E_{fwhm} value of $90.6/n$ mV will be obtained. However, as the concentration of the ferrocene adsorbates increases, the resulting voltammograms will become asymmetric, broadening, with an increase in the ΔE_p value.

The surface coverages were established from the CV profiles in 0.5 M H_2SO_4 using the relationship:

$$\Gamma = \frac{Q}{nFA} \quad (1)$$

where Q is the background-corrected charge under the cathodic or anodic waves, n = number of electrons involved in the redox process, F is the Faraday constant, and A is the area of the electrode. The surface coverage values are listed in Table 1. The average coverage increases as follows: SWCNT/FMCA (3.3×10^{-9} mol cm^{-2} , 2.0×10^{15} molecules cm^{-2}) > Cys-FMCA (1.6×10^{-9} mol cm^{-2} ,

Table 1 Summary of estimated voltammetric data obtained for the modified electrodes in 0.5 M H₂SO₄ (*n* = 4)

Electrodes	Electrochemical parameters				
	$E_{1/2}/\text{mV}$	$\Delta E_p/\text{mV}$	E_{fwhm}/mV	I_{pa}/I_{pc}	$10^9 \Gamma_a/\text{mol cm}^{-2}$
Au-Cys-FMCA	158 ± 2	4.8 ± 0.4	71 ± 3	0.8	1.6 ± 0.3
Au-Cys-FDCA	131 ± 3	14.7 ± 0.9	104 ± 2	0.9	1.1 ± 0.2
Au-Cys-SWCNT/FMCA	137 ± 2	95.2 ± 2.8	117 ± 5	1.1	3.3 ± 0.7
Au-Cys-SWCNT/FDCA	146 ± 4	4.7 ± 0.6	107 ± 3	1.1	1.4 ± 0.3

9.6×10^{14} molecules cm^{-2}) > SWCNT/FDCA (1.4×10^{-9} mol cm^{-2} , 8.4×10^{14} molecules cm^{-2}) > Cys-FDCA (1.1×10^{-9} mol cm^{-2} , 6.6×10^{14} molecules cm^{-2}). Since the maximum coverage expected of a monolayer of ferrocene is 4.5×10^{-10} mol cm^{-2} ,¹⁴ the coverages obtained in this work are between 3 and 8 times higher than expected. It should however be noted that other workers have observed similar or even higher values compared to our results. For example, Gooding *et al.*¹⁴ observed a much higher surface coverage ($1.8 \pm 0.9 \times 10^{-8}$ mol cm^{-2}) for randomly-distributed ferrocenes on SWCNTs and attributed this to the existence of a three-dimensional network of redox sites. Flavel *et al.*¹⁵ observed a 1.14×10^{-9} mol cm^{-2} and 9.59×10^{-8} mol cm^{-2} for macro- and nanoscale ferrocenemethanol-modified carbon nanotube electrodes on silicon, respectively. Chambers *et al.*¹² reported values of $\sim 4 \times 10^{14}$ molecules cm^{-2} . It is possible that the high surface coverages obtained in our experiments could be due to the long exposure time of cysteamine to gold electrode leading to more binding sites for the ferrocenes. The high surface coverage for the SWCNT/FMCA compared to SWCNT/FDCA suggests that the FMCA species probably coordinates more with SWCNTs *via* ester bonds as described above compared to its FDCA counterparts. It is known from Chidsey and co-workers^{2,36,37} that high concentration of ferrocene adsorbates leads to asymmetric and broad voltammograms, hence high ΔE_p values. Thus, the higher surface coverage shown by the FMCA-based electrode (Table 1) may further explain the poor ΔE_p and E_{fwhm} values of the SWCNT/FMCA compared to its SWCNT/FDCA counterpart.

Electrochemical impedance spectroscopy (EIS) is an important technique for probing heterogeneous electron transfer (HET) kinetics, especially at gold electrodes modified with self-assembled mono- or multi-layers of redox-active or redox-silent species.^{35,38–41} To establish the HET kinetics in this acidic solution, EIS experiments were carried out for each of the modified electrode, as described by Creager and Wooster.³⁸ Fig. 3a presents typical comparative Nyquist plots obtained for the four modified electrodes, biased at their approximate formal potential (~ 0.14 V vs. Ag|AgCl). Interestingly, the experimental data were satisfactorily fitted with the modified Randles electrical equivalent circuit (Fig. 3b), popularly used for modelling a redox-active monolayer on an electrode surface.^{35,38}

In the model, the R_s is the solution or electrolyte resistance, R_{ct} represents the electron-transfer resistance, while the true double layer capacitance (C_{dl}) is replaced by a constant phase element (CPE₁) and Warburg impedance (Z_w) replaced by another constant phase element (CPE₂ or capacitance of the

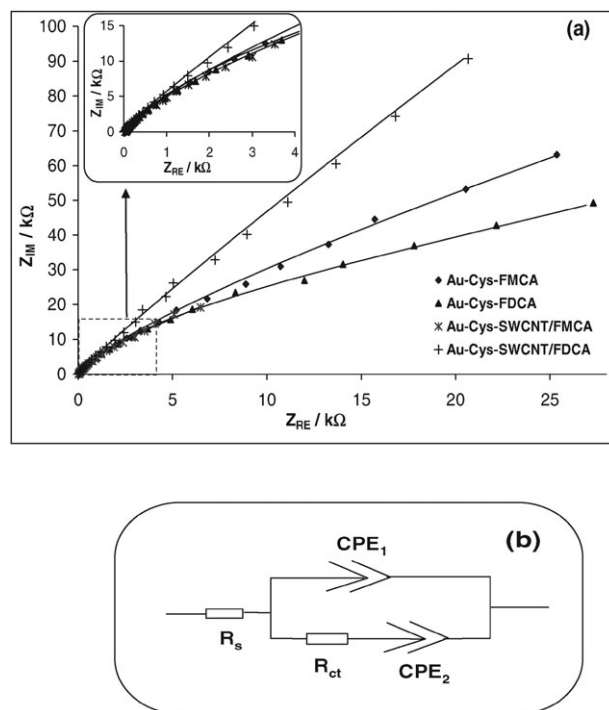


Fig. 3 Nyquist plot (a) and the (b) electrical equivalent circuit used to fit the impedance spectra of Au-Cys-FMCA, Au-Cys-FDCA, Au-Cys-SWCNT/FMCA and Au-Cys-SWCNT/FDCA electrodes obtained in 0.5M H₂SO₄. The symbols in (a) represent the experimental data, while solid lines are the fitted curves using the modified Randles equivalent circuit (b).

adsorbed molecules). The electron transfer rate constant (k_{et}) of each of the electrodes was obtained from:^{35,38}

$$k_{et} = \frac{1}{(2R_{ct}CPE_2)} \quad (2)$$

From Table 2, the k_{et} value increased as follows: Au-Cys-FMCA (~ 12 s⁻¹) \approx Au-Cys-SWCNT/FDCA (~ 10 s⁻¹) > Au-Cys-FDCA (~ 0.8 s⁻¹) > Au-Cys-SWCNT/FMCA (~ 0.3 s⁻¹), in close agreement with of the CV data listed in Table 1. It is interesting to observe how the neighbouring SWCNTs impact on the HET of the two different types of ferrocenes; SWCNTs improve the HET of ferrocene with exposed -COOH groups (FDCA) but decrease the HET of the ferrocene containing no pendant -COOH groups (FMCA).

The impedance of a CPE (Z_{CPE}) is defined as:^{42,43}

$$Z_{CPE} = \frac{1}{[Q(j\omega)^n]} \quad (3)$$

Table 2 Summary of estimated EIS parameters obtained for the modified electrodes in 0.5 M H₂SO₄

Electrodes	Electrochemical Impedometric parameters ^a						
	R_s/Ω	$CPE_1/\mu F$	n_1	$R_{ct}/k\Omega$	$CPE_2/\mu F$	n_2	k_{et}/s^{-1}
Au-Cys-FMCA	49.50 (0.72)	2.20 (3.57)	0.97 (0.44)	11.56 (27.61)	3.66 (1.59)	0.56 (1.09)	11.80
Au-Cys-FDCA	40.20 (0.66)	5.07 (1.08)	0.93 (0.18)	76.31 (6.20)	8.38 (2.40)	0.66 (3.46)	0.78
Au-Cys-SWCNT/FMCA	41.50 (1.50)	8.00 (1.86)	0.92 (0.36)	116.2 (6.36)	16.19 (5.23)	0.77 (2.76)	0.27
Au-Cys-SWCNT/FDCA	18.90 (1.16)	4.32 (2.81)	0.94 (0.71)	24.16 (29.22)	2.16 (5.02)	0.66 (1.27)	9.56

^a The values in brackets are the estimated error percentages obtained from the fitting using the circuit shown in Fig. 3b.

where Q is the frequency-independent constant relating to the surface electroactive properties, ω is the radial frequency, the exponent n arises from the slope of $\log Z$ vs. $\log f$, where f is the frequency (and has values $-1 \leq n \leq 1$). If $n = 0$, the CPE behaves as a pure resistor; for $n = 1$, CPE behaves as a pure capacitor; and for $n = -1$, CPE behaves as an inductor; while $n = 0.5$ corresponds to the Warburg impedance (Z_w) which is associated with the domain of mass transport control arising from the diffusion of ions to and from the electrode/solution interface. In general terms, CPE arises from such factors as (i) the nature of the electrode (*e.g.* roughness and polycrystallinity), (ii) distribution of the relaxation times due to heterogeneities existing at the electrode/electrolyte interface, (iii) porosity and (iv) dynamic disorder associated with diffusion.⁴⁴ Thus, the n_1 values in Table 2 are approximately close to 1.0 for an ideal capacitive behaviour, while the n_2 (that replaced the Warburg diffusion) are in the range between 0.56 and 0.77, describing the porous nature of the adsorbed film on the gold electrode. The apparent electron-transfer rate constant (k_{app}) values of the electrodes were obtained from eqn (1). The phase angles seen on the other Bode plots (*i.e.* $-\text{phase angle } (-\phi)$ vs. $\log f$, Fig. 4a) are in the range of 75–80°, which are less than the 90° expected of an ideal capacitive behaviour. The slopes of the Bode plots ($\log Z$ vs. $\log f$, Fig. 4b) are approximately similar (*ca.* -0.82 , $r^2 = 0.992$) at the mid frequency region, indicative of pseudo-capacitive behaviour. At high frequency regions, the slopes are almost zero, indicative of resistive behaviour at these high frequency regions.

Next, we interrogated the voltammetric behaviour of the electrodes when subjected to repetitive scanning in acidic electrolyte. This experiment is necessary for establishing the electrochemical stability of these redox-active adsorbates. Fig. 5 shows typical voltammograms (1st and 20th scans of

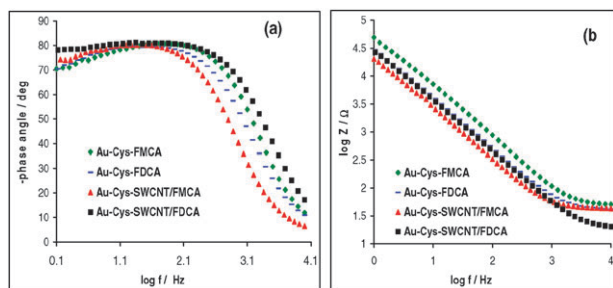


Fig. 4 Bode plots, phase angle vs. $\log f$ (a) and $\log Z$ vs. $\log f$ (b), of the impedance spectra of the modified electrodes in 0.5 M H₂SO₄. Experimental conditions are same as in Fig. 3.

each electrode) obtained on repetitively scanning each electrode in the 0.5 M H₂SO₄ solution. For all the electrodes, we observed that the first few scans (*ca.* 1–5 scans) remained unchanged. However, as the scan number was increased, the voltammograms became broadened (thereby increasing both the E_{fwhm} and ΔE_p values) stabilizing at about the 20th scan.

Unlike the low values of the ΔE_p seen at the first five scans (Table 1), at the 20th scan the ΔE_p values increased to ~ 100 mV for the Cys-FMCA, > 200 mV for the Cys-FDCA, and ~ 100 mV for the SWCNT/FDCA. For the SWCNT/FMCA, however, there was a slight decrease from its original 95 mV to about 80 mV at the 20th scan. This is an interesting observation considering that such behaviour has only been observed when the concentration of ferrocene adsorbates is increased.^{2,36,37} Besides, Finklea⁴⁵ had demonstrated that the cause of such behavior arising from an increase in concentration of the ferrocene to be due to double-layer effects which arise from a rapid increase in surface charge during oxidation. In other words, surface charge effectively decreases the potential difference between the remaining unoxidized ferrocenes and the surface, thus requiring a high applied bias for the oxidation of these ferrocene. However, in our case, since the concentration of the ferrocene species in each electrode is the same at all scans, the assumption of the asymmetric and broadening nature of the redox peaks during repetitive scanning may probably be due to the disordering of the initial well-separated ferrocene species to a form where they now interact with one another. Simply put, during the first few scans of the

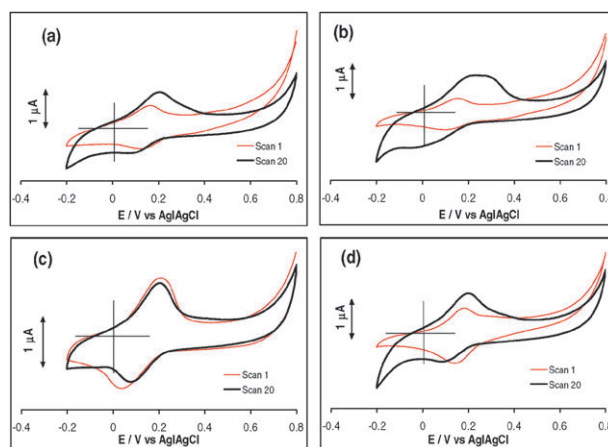


Fig. 5 Comparative cyclic voltammetric evolutions of the modified gold electrode in 0.5M H₂SO₄ obtained at the 1st and the 20th scans for (a) Au-Cys-SWCNT/FDCA, (b) Au-Cys-FDCA, (c) Au-Cys-SWCNT/FMCA and (d) Au-Cys-FMCA. Scan rate = 25 mV s⁻¹.

ferrocene-based electrodes, the ferrocenes head groups are still well-separated and do not interact with one another (indicated by the near-ideal E_{fwhm} value, Table 1). However, as the scanning number is increased, these ferrocenes may now begin to interact with one another and/or the double layer effects ensue. Also, it is worth noting from Fig. 3 that without the SWCNTs, the ΔE_p value is higher compared to when ferrocene is co-assembled with the SWCNTs, suggesting that SWCNTs tend to suppress this double-layer effects (for example, compare the CVs of Au-Cys-FMCA and Au-Cys-SWCNT/FMCA).

The existence of the ferrocene redox activity during the repetitive scanning indicates that the modified electrodes exhibited strong electrochemical stability in 0.5 M H_2SO_4 solution. Such stability is important for their electrochemical studies as well as their potential applications in acidic pH conditions. Thus, all subsequent experiments with each of the ferrocene-based electrodes were performed after 20 continuous cyclic voltammetric scanning in 0.5 M H_2SO_4 solution at 50 $mV s^{-1}$.

3.3 Heterogeneous electron transfer dynamics in a redox probe, $[Fe(CN)_6]^{4-}/[Fe(CN)_6]^{3-}$

Electron transport properties of the electrodes were studied in PBS solution containing the outer-sphere redox probe, $[Fe(CN)_6]^{4-}/[Fe(CN)_6]^{3-}$ (pH 7.2). Typical comparative CVs are exemplified in Fig. 6. The ΔE_p values for all the electrodes are essentially the same (~ 110 mV) with the same equilibrium potential ($E_{1/2} \approx 100$ mV). Since electrochemical impedance spectroscopy (EIS) provides a more detailed description of an electrochemical system⁴⁶ than cyclic voltammetry does, EIS was used to follow the electron transfer kinetics occurring at these electrodes.

The EIS measurements were performed at the formal potential ($E_{1/2} \approx 100$ mV). Fig. 7 shows examples of the Nyquist plots obtained for all the electrodes in the outer-sphere redox probe, $[Fe(CN)_6]^{4-}/[Fe(CN)_6]^{3-}$ solution (pH 7.2).

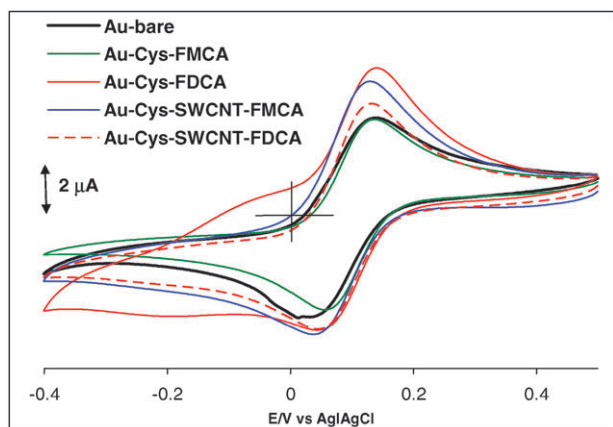


Fig. 6 Comparative cyclic voltammograms of the bare and modified gold electrodes obtained in 0.1 M $[Fe(CN)_6]^{4-}/[Fe(CN)_6]^{3-}$ (PBS, pH 7.2); bare Au, Au-Cys-FDCA, Au-Cys-FMCA, Au-Cys-SWCNT/FDCA and Au-Cys-SWCNT/FMCA. Scan rate = 25 $mV s^{-1}$.

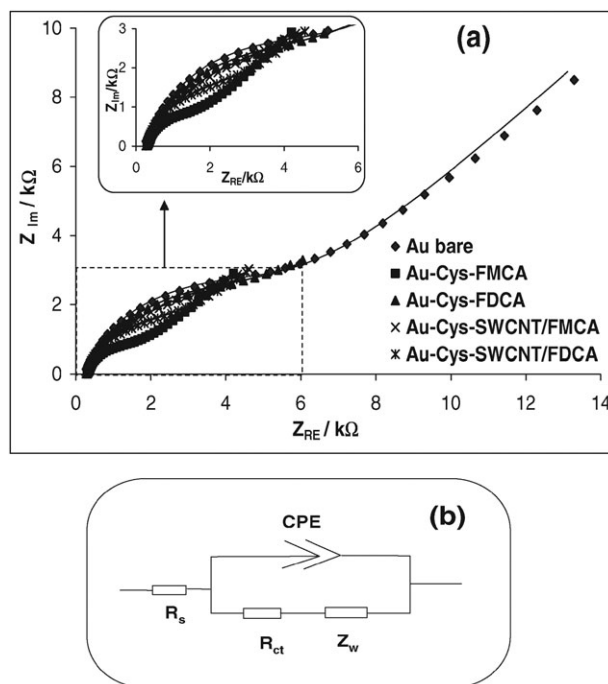


Fig. 7 Nyquist plot (a) and the electrical equivalent circuit (b) used to fit the impedance spectra of bare Au, Au-Cys-FMCA, Au-Cys-FDCA, Au-Cys-SWCNT/FMCA and Au-Cys-SWCNT/FDCA obtained in 0.1 M $[Fe(CN)_6]^{4-}/[Fe(CN)_6]^{3-}$ (PBS, pH 7.2). The symbols in (a) represent the experimental data, while solid lines are the fitted curves using the modified Randles equivalent circuit (b).

The EIS data were satisfactorily fitted with the modified Randles equivalent circuit model (Fig. 7b), judged by the relative % errors (Table 2) wherein the true capacitance is replaced by the constant phase element (CPE). In this model the R_s is the solution/electrolyte resistance, R_{ct} represents the electron-transfer resistance, while the Z_w is the Warburg impedance. The apparent electron transfer rate constant (k_{app}) of each of the electrodes was obtained from eqn (4):

$$k_{app} = \frac{RT}{n^2 F^2 A R_{ct} C} \quad (4)$$

where n is the number of electrons transferred (1), F is the Faraday constant, R is the ideal gas constant, T is the Kelvin temperature, A is the experimentally-determined area of the electrode, the R_{ct} value is obtained from the fitted Nyquist plots, C is the concentration of the $[Fe(CN)_6]^{3-}$ (in $mol cm^{-3}$, the concentration of $[Fe(CN)_6]^{3-}$ and $[Fe(CN)_6]^{4-}$ are equal). From Table 3, the k_{app} values increase as follows: Au-Cys-FMCA > Au-Cys-SWCNT/FDCA > Au-Cys-SWCNT/FMCA \approx Au-Cys-FDCA > bare Au. Within the limits of experimental errors, the trend in the electron transport kinetics seen in the redox probe ($[Fe(CN)_6]^{3-}/[Fe(CN)_6]^{4-}$ solution is somewhat comparable with that observed with the H_2SO_4 solution. However, we should not expect them to necessarily follow the same trend. This is because, first, the experiments were performed at different electrolyte conditions and, second, unlike the experiment in the H_2SO_4 which provides insight into the redox processes (Fe^{3+}/Fe^{2+}) of the electrode-confined ferrocene

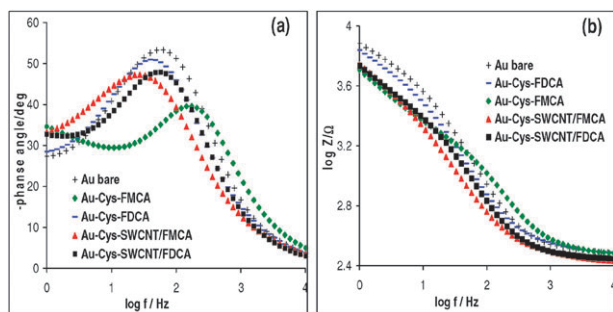


Fig. 8 Bode plots, (a) phase angle vs. $\log f$ and (b) $\log Z$ vs. $\log f$ of the impedance spectra of the modified electrodes in redox probe ($[\text{Fe}(\text{CN})_6]^{4-}/[\text{Fe}(\text{CN})_6]^{3-}$) PBS solution (pH 7.2). Experimental conditions are same as in Fig. 7.

species, the experiment with the $[\text{Fe}(\text{CN})_6]^{3-}/[\text{Fe}(\text{CN})_6]^{4-}$ solution essentially interrogates the extent to which the immobilized ferrocene species permit the permeation of the redox probe and/or enhance the faradaic response of the $[\text{Fe}(\text{CN})_6]^{3-}$ and $[\text{Fe}(\text{CN})_6]^{4-}$ species. The n values (Table 3) are in the range 0.76 and 0.86, indicative of pseudocapacitive behaviour. From the Bode plots (Fig. 8), the values of the slopes of the $\log Z$ vs. $\log f$ plot (*ca.* -0.57) as well as the phase angles (less than 90°) confirm the pseudocapacitive behaviour. Notice from the Bode plot (Fig. 8a) that the peak of the Au-Cys-FMCA occurred at a higher frequency (~ 200 Hz) than the others which occurred at frequencies < 100 Hz. This indicates that the electrochemical reaction rates is faster with Au-Cys-FMCA, as f usually represents the time constants of an electrochemical reaction.

3.4 Electrocatalytic detection of thiocyanate

Having established that the neighbouring SWCNTs impact on the HET of the ferrocene molecular assemblies, we thought it was necessary to establish the extent to which these SWCNTs could impact on the electrocatalytic behaviour of the ferrocenes. For this study, we chose thiocyanate (SCN^-) as a model analyte. Fig. 9 shows the comparative square wave voltammetric (SWV) evolutions obtained at constant concentration (1 mM) of SCN^- at the various electrodes in PBS (pH 4.8). This pH condition was chosen for this experiment as it is well known to enhance the detection of thiocyanate.^{47–49}

The catalytic behaviour (in terms of onset potential) follows this trend: bare Au (0.64 V) > Au-Cys-SWCNT (0.62 V) > Au-Cys-FMCA \approx Au-Cys-FDCA (0.48 V) > Au-Cys-SWCNT/FMCA (0.50 V) > Au-Cys-SWCNT/FDCA (0.42 V). Au-Cys-SWCNT/FDCA and Au-Cys-SWCNT/FMCA gave

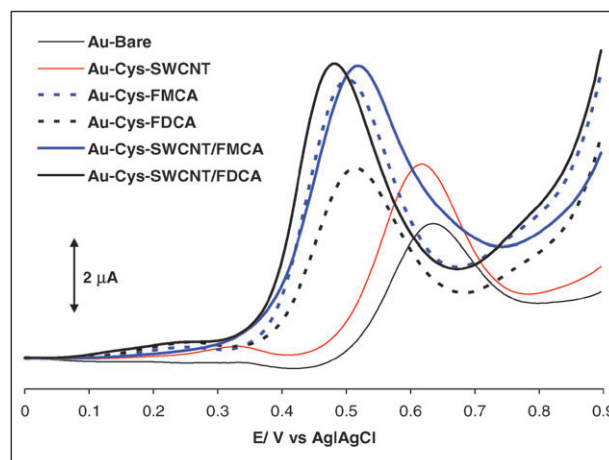


Fig. 9 Comparative square wave voltammograms of the bare gold and modified gold electrodes obtained in PBS solution (pH 4.7) containing 1 mM SCN^- .

the highest current response with less positive onset potential (0.3 V) compared to other electrodes. However, Au-Cys-SWCNT/FDCA recorded the least peak potential for the oxidation of SCN^- at 0.48 V, possibly due to the repulsive interaction between the negatively charged SCN^- and high number of surface $-\text{COOH}$ species at the FDCA. Second, the ferrocene-based electrode without the SWCNTs (*i.e.* Au-Cys-FDCA and Au-Cys-FMCA) exhibited a more positive onset potential (~ 0.35 V) for the oxidation of SCN^- than the ones containing SWCNTs (Au-Cys-SWCNT/FDCA and Au-Cys-SWCNT/FMCA) at 0.3 V. This improved response towards the detection of SCN^- species is due the combined synergistic activities of good electrocatalysts (FDCA or FCA) and the efficient electronic conducting nanowires (SWCNTs). This finding is quite remarkable especially when compared to data previously obtained at the iron (ii) phthalocyanine (FePc) and SWCNT-iron (ii) octa(hydroethylthio) phthalocyanine (FeOHETPc) SAM modified electrodes where the oxidation of SCN^- occurred at peak potentials ≥ 0.60 V.^{48,49} Also, the current responses recorded in this present work for the Au-Cys-SWCNT/FDCA and Au-Cys-SWCNT/FMCA for the same concentration of SCN^- (1 mM) is approximately twice that reported for the FeOHETPc and FePc-SAM modified electrodes.^{48,49}

Chronoamperometric detection of SCN^- at the Au-Cys-SWCNT/FDCA (Fig. 10) gave a higher sensitivity of $\sim 7.5 \times 10^{-3} \text{ AM}^{-1}$ compared to the $4.2 \times 10^{-3} \text{ AM}^{-1}$ obtained before.²⁰ The detection limit was of $\sim 1 \mu\text{M}$. We

Table 3 Summary of estimated EIS parameters obtained in 0.1 M $[\text{Fe}(\text{CN})_6]^{4-}/[\text{Fe}(\text{CN})_6]^{3-}$ (PBS, pH 7.2)

Electrodes	Electrochemical impedemetric parameters ^a					
	R_s/Ω	CPE/ μF	n	$R_{ct}/\text{k}\Omega$	$Z_w/\mu\Omega$	$10^3 k_{app}/\text{cm s}^{-1}$
Au	305.81 (0.43)	5.32 (2.02)	0.86 (0.37)	5.12 (1.15)	104.0 (0.90)	1.93 ± 0.02
Au-Cys-FMCA	300.90 (0.40)	0.44 (4.16)	0.86 (0.58)	1.41 (1.64)	98.3 (0.68)	6.98 ± 0.11
Au-Cys-FDCA	310.00 (0.63)	0.63 (5.56)	0.82 (0.69)	4.93 (3.03)	118.7 (4.52)	2.00 ± 0.06
Au-Cys-SWCNT/FMCA	271.70 (0.53)	0.52 (5.59)	0.76 (0.68)	4.65 (3.87)	126.6 (4.67)	2.12 ± 0.08
Au-Cys-SWCNT/FDCA	280.10 (0.33)	0.98 (3.22)	0.86 (0.45)	2.58 (1.83)	107.6 (1.28)	3.82 ± 0.07

^a The values in brackets are the estimated error percentages obtained from the fitting using the circuit shown in Fig. 7b.

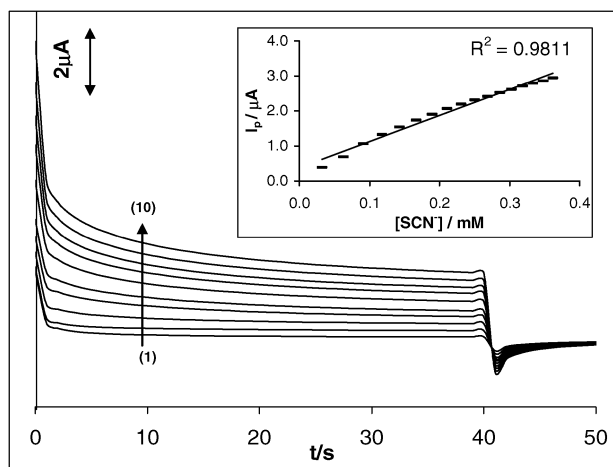


Fig. 10 Typical double potential step chronoamperometric transients at Au-Cys-SWCNT-FDCA in PBS solution (pH 4.7) following the addition of thiocyanate (numbers 1–10 correspond to the increasing concentrations). The potential was stepped at +0.45 and +0.20 V. Inset is the plot of the chronoamperometric current vs. $[\text{SCN}^-]$.

also tested the suitability of the ITO-nanoAu-Cys-SWCNT/FDCA for a possible application as a disposable, one-shot electrode system for a quick chronoamperometric detection of SCN^- (not shown). The ITO-nanoAu used for this experiment was fabricated using the mediated growth mechanism as reported in a previous communication.⁵⁰ With this electrode, we also recorded chronoamperometric sensitivity of 7.3×10^{-3} and a limit of detection of ~ 13 nM for the SCN^- , comparable to the data obtained at the modified bulk gold electrode.

4. Conclusions

This study has for the first time interrogated the electrochemistry of electron transfer dynamics of ferrocene-terminated self-assembled monolayers (SAMs), co-adsorbed with SWCNTs on a gold electrode. Factors influencing electron transport within organized molecular assemblies are crucial for the potential applications of such platforms in many areas such as in molecular electronics, chemical and biological sensings. The important findings in this work should be emphasized. First, the neighbouring SWCNTs in the ferrocene molecular assemblies exert distinct impacts on the global electron transport and electrocatalytic behaviour of the ferrocenes. Second, that the presence of SWCNTs in the ferrocene assembly synergistically enhances the electrocatalytic detection of thiocyanate compared to the ferrocene or SWCNTs alone. In a nutshell, this study has provided some useful physical insights into the impact of local SWCNTs microenvironment surrounding a redox-active and electrocatalytic molecule (not only ferrocene but other related organometallic species) on the dynamics of electron transport between solution species and electrode. We envisage that these insights have provided some thoughts on the factors that must be considered when designing molecular-scale electronics or electrocatalytic devices that employ CNT and ferrocenes (or related species).

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